

(4) Assertion is wrong but reason is correct statement.

Ans. (2)

Sol. Due to half filled orbital configuration nitrogen has more ionisation energy than oxygen.

2. Column-I

- (i) Siderite
- (ii) Malachite
- (iii) Carnalite
- (iv) Calamine

Column-II

- (a)  $KCl \cdot MgCl_2 \cdot 6H_2O$
- (b)  $CuCO_3 \cdot Cu(OH)_2$
- (c)  $ZnCO_3$
- (d)  $FeCO_3$

Correct match is :

	I	II	III	IV		I	II	III	IV
(1)	d	b	a	c	(2)	a	b	c	d
(3)	d	a	b	c	(4)	d	b	c	a

Ans. (1)

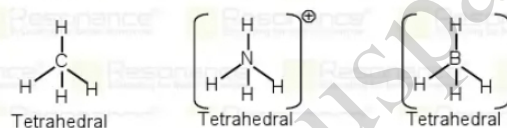
Sol. Siderite  $\Rightarrow FeCO_3$   
 Malachite  $\Rightarrow CuCO_3 \cdot Cu(OH)_2$   
 Carnalite  $\Rightarrow KCl \cdot MgCl_2 \cdot 6H_2O$   
 Calamine  $\Rightarrow ZnCO_3$

3. Which of the following is correct option regarding  $CH_4$ ,  $NH_4^+$  and  $BH_4^-$

- (1) All are isoelectronic & tetrahedral.
- (2) All are not isoelectronic but tetrahedral.
- (3) All are isoelectronic but only two are tetrahedral.
- (4) All are isoelectronic but all are not tetrahedral.

Ans. (1)

Sol. Species  $CH_4$   $NH_4^+$   $BH_4^-$   
 No. of electron 10 10 10

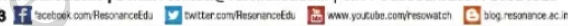
Structure   
 Tetrahedral Tetrahedral Tetrahedral

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4. Which set of compounds contain carbonate ion ?

- (1) Baking soda, Washing soda
- (2) Baking soda, Caustic soda
- (3) Washing soda, Caustic soda
- (4) Only Washing soda

Ans. (1)

Sol. Compound Formula

- (1) Baking soda  $NaHCO_3$
- (2) Washing soda  $Na_2CO_3 \cdot 10H_2O$
- (3) Caustic soda  $NaOH$

5. 1 Mole of  $CoCl_3 \cdot xNH_3$  on reaction with excess of  $AgNO_3$  give 2 moles of  $AgCl$  then value of X is :

Ans. (5)

Sol.  $CoCl_3 \cdot xNH_3 + AgNO_3$  (excess)  $\longrightarrow AgCl$

1 mole 2 moles

It means 2 Cl are outside the co-ordinations sphere & co-ordination number of Co is 6

So possible complex is  $[Co(NH_3)_5Cl]Cl_2$

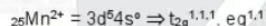
so x = 5

6. The magnetic moment (spin only) of complex  $[MnBr_6]^{4-}$  is ..... BM

[Report your answer to nearest integer]

Ans. (6)

Sol.  $[\text{MnBr}_6]^{4-}$



number of unpaired electrons = 5

$$\mu(\text{spin only}) = \sqrt{n(n+2)} \text{ BM} = \sqrt{5(5+2)} \text{ BM} = \sqrt{35} \text{ BM} = 5.916 \text{ BM} \approx 6 \text{ BM}$$

7. For a first order reaction  $K = 6.3 \times 10^{-18} e^{-2000/T}$ . Then the value of activation energy in KJ is :

[given  $R = 8.314 \text{ J/mole} \times \text{K}$ ]

[Report your answer to nearest integer]

Ans. (216)

Sol.  $K = Ae^{-(E_a/R)/T}$

$$K = 6.3 \times 10^{-18} e^{-2000/T}$$

$$\frac{E_a}{R} = 26000 = 26 \times 10^3$$

$$E_a = 26 \times 10^3 \times 8.314 = 216.164 \times 10^3 \text{ J} = 216.164 \text{ KJ}$$

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8. In 100 L vessel at 610 K, 4 mole of Ar and 5 mole of  $\text{PCl}_5$  are taken. At equilibrium total pressure of gases is 6 atm, then value of  $K_p$  is :

$$[R = 0.082 \frac{\text{atm} \times \text{L}}{\text{mole} \times \text{K}}]$$

[Report your answer to nearest integer]

Ans. (2)

Sol.  $\text{PCl}_5(\text{g}) \rightleftharpoons \text{PCl}_3(\text{g}) + \text{Cl}_2(\text{g})$

5 mole            0            0

(5 - x)            x            x

Total moles at equilibrium = (5 + x) +  $n_{\text{Ar}}$  = (5 + x) + 4 = (9 + x)

$$n_{\text{total}} = \frac{PV}{RT} = \frac{6 \times 100}{0.082 \times 610} = 11.995 = 12 \text{ moles}$$

$$9 + x = 12 \text{ moles}$$

$$x = 3 \text{ moles}$$

$$\text{Pressure of } (\text{PCl}_5 + \text{PCl}_3 + \text{Cl}_2) = \frac{8}{12} \times 6 = 4 \text{ atm}$$

$$K_p = \frac{P_{\text{PCl}_3} \times P_{\text{Cl}_2}}{P_{\text{PCl}_5}} = \frac{\left(\frac{3}{8} \times 4\right) \left(\frac{3}{8} \times 4\right)}{\left(\frac{2}{8} \times 4\right)} = \left(\frac{3}{2}\right) \left(\frac{3}{2}\right) = \left(\frac{9}{4}\right) = 2.25$$

9. For a solution containing non volatile non electrolyte solute concentration is 1.5 m. The elevation in boiling point ( $\Delta T_b$ ) is 4 K while at concentration of 4.5 m depression in freezing point ( $\Delta T_f$ ) is 4 K then find ratio of  $\left(\frac{K_b}{K_f}\right)$ .

Ans. (3)

Sol.  $\Delta T_b = K_b \times m_1$

$$\Delta T_f = K_f \times m_2$$

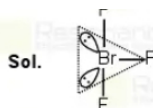
$$\Rightarrow \frac{\Delta T_b}{\Delta T_f} = \frac{K_b \times 1.5}{K_f \times 4.5} = \frac{4 \text{ K}}{4 \text{ K}}$$

$$\frac{K_b}{K_f} = 3$$

10. Shape and number of lone pair electrons in  $\text{BrF}_3$  is :

(1) Bent T-shape, 2      (2) Bent T-shape, 1      (3) See-Saw, 2      (4) See-Saw, 1

Ans. (1)



Bent T-shape with Two unpaired electron.

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11. 42.14 % (w/v) NaCl solution is used to coagulate 1 L of colloidal solution in 10 hours, then coagulation value for 2 hours is :

- (1) 36 mole                      (2) 36 millimole                      (3) 1440 mole                      (4) 1440 millimole

Ans. (4)

Sol. Coagulation value =  $\frac{\text{millimoles of electrolyte}}{\text{Volume of solution in L}}$

$$\text{Molarity of NaCl} = \frac{\% (w/v) \times 10}{\text{GMM}} = \left[ \frac{42.14 \times 10}{58.5} \right] = 7.2 \text{ M}$$

millimole of NaCl electrolyte = 7.2 mole = 72000 millimole

Coagulation value for 10 hours =  $\frac{\text{millimoles of electrolyte}}{\text{Volume of solution in L}} = 72000$

For 2 hours Coagulation value =  $\left( \frac{72000 \times 2}{10} \right) = 1440 \text{ millimole}$

12. For the following cell: Pt(s) | H<sub>2</sub>(g) | H<sup>+</sup>(aq) || Cu<sup>2+</sup>(aq) | Cu at pH = 3, E<sub>cell</sub> = 0.31 V and [Cu<sup>2+</sup>] = 10<sup>-x</sup>, then value of x is .... [Given E<sup>o</sup><sub>Cu<sup>2+</sup>/Cu</sub> = 0.34 V]

Ans. (7)

Sol. Anode : H<sub>2</sub>(g) → 2H<sup>+</sup> + 2e<sup>-</sup>

Cathode : Cu<sup>2+</sup> + 2e<sup>-</sup> → Cu(s)

Overall : H<sub>2</sub>(g) + Cu<sup>2+</sup>(aq) → 2H<sup>+</sup>(aq) + Cu(s)

$$E_{\text{cell}} = E^{\circ}_{\text{cell}} - \frac{0.059}{2} \log \frac{[\text{H}^+]^2}{[\text{Cu}^{2+}]}$$

$$0.31 = 0.34 - \frac{0.06}{2} \log \left( \frac{[\text{H}^+]^2}{[\text{Cu}^{2+}]} \right)$$

$$0.31 = 0.34 + 0.03 [-\log [\text{H}^+]^2 + \log [\text{Cu}^{2+}]]$$

$$0.31 = 0.34 + 0.03 [2\text{pH} + \log [\text{Cu}^{2+}]]$$

$$-0.03 = 0.03 [2\text{pH} + \log [\text{Cu}^{2+}]]$$

$$-1 = 6 + \log [\text{Cu}^{2+}]$$

$$-7 = \log [\text{Cu}^{2+}]$$

$$\log [\text{Cu}^{2+}] = \log 10^{-7}$$

$$[\text{Cu}^{2+}] = 10^{-7}$$

$$x = 7$$

13. Statement-I : CuSO<sub>4</sub>·5H<sub>2</sub>O contain Cu-O Bond.

Statement-II : Sulphur and oxygen donate it's electron pair and act as ligand.

- (1) Both S<sub>1</sub> & S<sub>2</sub> are true                      (2) Both S<sub>1</sub> & S<sub>2</sub> are false  
(3) S<sub>1</sub> is true & S<sub>2</sub> is false                      (4) S<sub>1</sub> is false & S<sub>2</sub> is true

Ans. (3)

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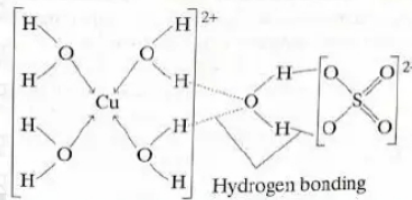
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Sol.  $\text{CuSO}_4 \cdot 5\text{H}_2\text{O} \Rightarrow [\text{Cu}(\text{H}_2\text{O})_4]\text{SO}_4 \cdot \text{H}_2\text{O}$



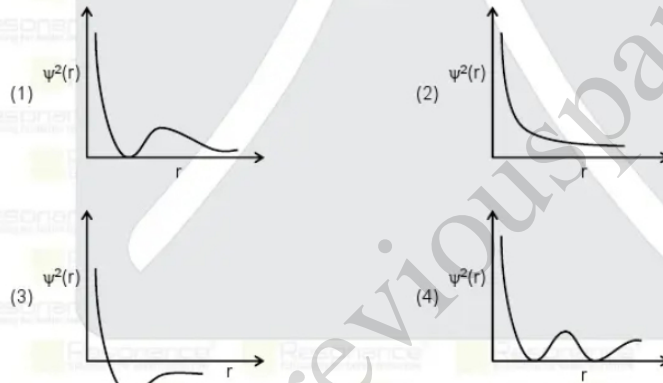
14. An inorganic Compound on reaction with  $\text{BaCl}_2$  give white ppt which on reaction with dilute  $\text{HCl}$  which on reaction with dilute  $\text{HCl}$  give characteristics smell. Which anion is present in inorganic compound

- (1)  $\text{I}^-$  (2)  $\text{S}^{2-}$  (3)  $\text{SO}_3^{2-}$  (4)  $\text{SO}_4^{2-}$

Ans. (3)

Sol.  $\text{SO}_3^{2-} + \text{Ba}^{2+} \longrightarrow \text{BaSO}_3 \downarrow$  (White ppt)  $\xrightarrow{\text{dil HCl}}$   $\text{SO}_2(\text{g}) \uparrow$  (Burning sulphur like smell)  
 $\text{S}^{2-} + \text{Ba}^{2+} \longrightarrow$  No PPT

15. Identify the correct graph for 2s-orbital for  $\psi^2(r)$  vs  $r$ .



Ans. (1)

16. A container contain 128 gram  $\text{O}_2$  (g) and 16 gram  $\text{H}_2$ , then volume of gaseous mixture at STP is (Report your answer to nearest integer)

Ans. (269)

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Sol.  $n_{\text{O}_2} = \frac{128}{32} = 4$  mole

$n_{\text{H}_2} = \frac{16}{2} = 8$  mole

$n_{\text{Total}} = 12$

Volume at STP =  $12 \times 22.4 = 268.8 \text{ L} \approx 269 \text{ L}$

17. What is the value of x in :

0.002858 × 0.112 ...

$$\frac{0.5702}{0.5702} = x \quad (1) 0.00056 \quad (2) 0.000561 \quad (3) 0.000503 \quad (4) 0.0006$$

Ans. (2)

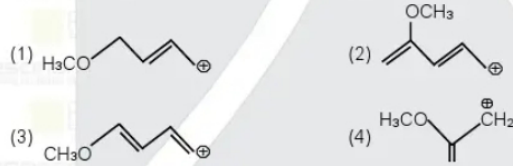
Sol.  $\frac{0.00285 \times 0.112}{0.5702} = \frac{0.0003200}{0.5702} = 0.000561$

18. A compound decompose according to 1<sup>st</sup> order reaction, then find time taken (in hours) to reduce concentration from initial value to 6.25 % if its half life is 5 hour.

Ans. (20)

Sol.  $100 \xrightarrow{t_{1/2}} 50 \xrightarrow{t_{1/2}} 25 \xrightarrow{t_{1/2}} 12.5 \xrightarrow{t_{1/2}} 6.25$   
total time =  $4T_{1/2} = 20$  hours

19. Identify most stable carbocation out of following.



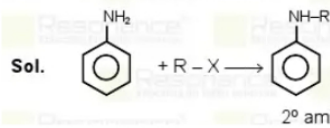
Ans. (3)

Sol. Due to extend conjugation most stable carbocation is

20. Friedal craft alkylation of aniline gives

- (1) Secondary amine (2) Amide product after attack on aniline  
(3) ortho/para alkyl derivative (4) Meta alkyl derivative

Ans. (1)



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21. On heating which structure not affected.

- (1) Secondary structure of protein (2) Primary structure of protein  
(3) Tertiary structure of protein (4) Quaternary structure of protein

Ans. (2)

Sol. During denaturation of protein 2 and 3 structure are destroyed but 1<sup>o</sup> structure remain intact.

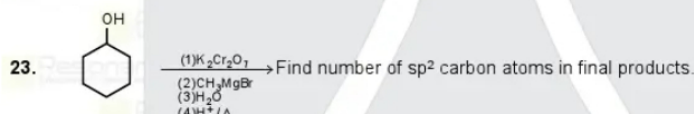
22. **Statement-I** : Dacron is an example of polyester compound.

**Statement-II** : Dacron is a combination of terphthalic acid & ethylene glycol.

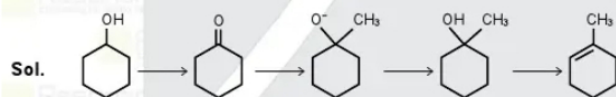
- (1) Statement-I and statement-II both are correct and statement-II is correct explanation of statement-I.  
(2) Statement-I and statement-II both are correct statements but statement-II is not correct explanation of statement-I.  
(3) statement-I is correct but statement-II is wrong statement.  
(4) statement-I is wrong but statement-II is correct statement

Ans. (1)

Sol. It is fact.



Ans. 2.0



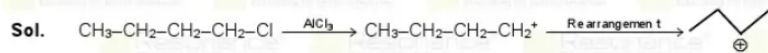
No. of sp<sup>2</sup> carbon are 2 in final product.



Most stable carbocation possible in above reaction is :



Ans. (1)



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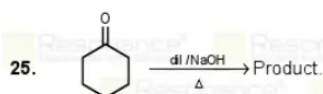
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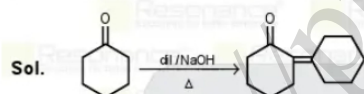
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Find out number of π-bonds in product from by above reaction.

Ans. (2)



26. Chloroxylenol and terpineol work as :

- (1) Antiseptic (2) Disinfectant (3) Antipyretic (4) Antibiotic

Ans. (1)

Sol. Commonly used antiseptic Dettol is a mixture of chloroxylenol and terpineol.

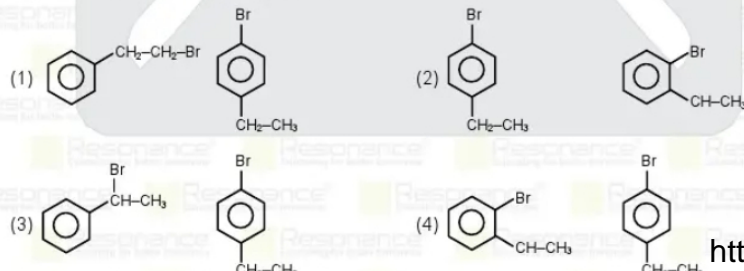
27. Reaction involve in troposphere during acid rain.

- (1)  $\text{H}_2\text{S} + \text{O}_2 \longrightarrow \text{S} + \text{H}_2\text{O}$  (2)  $\text{S} + \text{NaOH} \longrightarrow \text{Na}_2\text{S} + \text{Na}_2\text{S}_2\text{O}_3 + \text{H}_2\text{O}$   
 (3)  $\text{I}_2 + \text{Na}_2\text{S}_2\text{O}_3 \longrightarrow \text{Na}_2\text{S}_4\text{O}_6 + \text{NaI}$  (4)  $2\text{SO}_2 + \text{O}_2 + 2\text{H}_2\text{O} \longrightarrow 2\text{H}_2\text{SO}_4$

Ans. (4)

Sol. SO<sub>2</sub> and NO<sub>2</sub> after oxidation and reaction with H<sub>2</sub>O are major contributors to acid rain.

28. 184 g per mole of given compound having C = 52.4%, H = 4.9% and Br = 42.7% and both A and B react with KMnO<sub>4</sub> & will give benzoic acid and para-bromo benzoic acid respectively than identify compound A and B. Compound A is optically active.



Ans. (3)

<https://previouspaper.in>

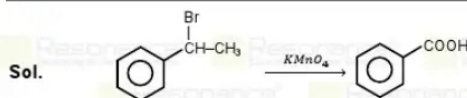
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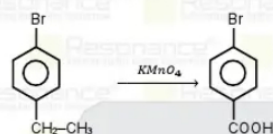
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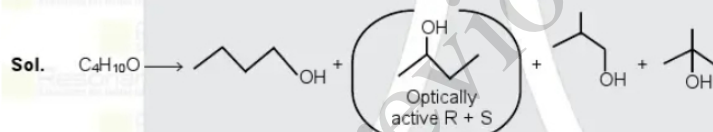
Optically active (A)



(B)

29. Find out number of chiral alcohol of molecular formula  $C_4H_{10}O$ .

Ans. (2)



Only 2-Butanol is chiral with R or S configuration.


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