

2. **Statement-I:**  $O^{2-}$  and  $Mg^{2+}$  have equal radii

**Statement-II:** They are isoelectric

- (1) Both Statement-I & Statement-II are True.  
 (2) Both Statement-I & Statement-II are False.  
 (3) Statement-I is True while Statement-II is False.  
 (4) Statement-I is False while Statement-II is True.

**Ans.** (4)

Sol.	ion	$O^{2-}$	$Mg^{2+}$
	No. of $e^-$	10	10
	Z	8	12

Order of ionic size  $\Rightarrow O^{2-} > Mg^{2+}$

3. Which of the following complex have maximum value of crystal field splitting energy.

- (1)  $[Co(CN)_6]^{3-}$  (2)  $[Cu(NH_3)_4]^{2+}$  (3)  $[Co(H_2O)_6]^{2+}$  (4)  $[Ti(H_2O)_6]^{3+}$

**Ans.** (2)

Sol.	Complex	Hybridisation
(1)	$[Co(CN)_6]^{3-}$	$d^2sp^3$
(2)	$[Cu(CN)_4]^{2+}$	$dsp^2$
(3)	$[Co(H_2O)_6]^{2+}$	$sp^3d^2$
(4)	$[Ti(H_2O)_6]^{3+}$	$d^2sp^3$

$\Delta_{sp} = 1.3 \Delta_0$

4. When acidified  $KMnO_4$  reacts with oxalic acid gives a manganese product. Find the magnetic moment (spin only) of Mn in this product.

(Report your answer to nearest integer)

**Ans.** (6)



$_{25}Mn^{2+} = 3d^5$

No. of unpaired electron = 5

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$$\begin{aligned} \mu(\text{spin only}) &= \sqrt{n(n+2)} BM \\ &= \sqrt{5(5+2)} \\ &= \sqrt{35} \\ &= 5.916 \end{aligned}$$

**Ans.** 6

5. 2 gram of non-volatile solute dissolve in each 200 gram of two different solvents A and B and have same molality. It ratio of  $K_b$  of solvent A and B is 1 : 8 & the ratio of elevation in boiling point is x : y, then value of y is \_\_\_\_\_.

**Ans.** (8)

**Sol.**  $\Delta T_b = K_b \times m$

$$\frac{(\Delta T_b)_I}{(\Delta T_b)_{II}} = \frac{(K_b)_I}{(K_b)_{II}} = \frac{1}{8} = \frac{x}{y}$$

So y = 8

6. **S<sub>1</sub>** : At 20°C molality of KCl solution is 'X' m, on decreasing temperature to -10°C molality remain unchanged.

**S<sub>2</sub>** : On changing temperature mass remain unchanged.

- (1) Both  $S_1$  &  $S_2$  are true (2) Both  $S_1$  &  $S_2$  are false

(3)  $S_1$  is true &  $S_2$  is false (4)  $S_1$  is false &  $S_2$  is true

**Ans.** (1)

**Sol.** Molality & mass are temperature independent so on changing temperature molality & mass remain unchanged.

7. What product are obtained in following reaction,  $H_2SO_4 + BaO_2 \rightarrow$  Product.

(1)  $H_2 + O_2$  (2)  $H_2O$  (3)  $BaSO_4 + H_2O_2$  (4)  $H_2SO_8 + H_2O_2$

**Ans.** (3)

**Sol.**  $BaO_2 \cdot 8H_2O(s) + H_2SO_4(aq) \rightarrow BaSO_4(s) + H_2O_2(aq) + 8H_2O(l)$

8. Two element A & B form two different compound  $A_2B$  &  $AB_3$ . If 0.15 mole of both compound have equal weight then what is the ratio of molar mass of A & B.

**Ans.** (2)

**Sol.** Let molar mass of A is a

B is b

again  $\Rightarrow 0.15 [2a + b] = 0.15 [a + 3b]$

$$a = 2b \Rightarrow \left(\frac{a}{b}\right) = \frac{2}{1}$$

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9. Match the following.

**List-I**

(i)  $Si(CH_3)_4$

(ii)  $Si(CH_3)_2(OH)_2$

(iii)  $Si(CH_3)(OH)_3$

(iv)  $Si(CH_3)_3(OH)$

**List-II**

(a) Silane

(b) 2D silicone

(c) Chain silicone

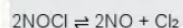
(d) Dimeric silicone

	I	II	III	IV
(1)	a	c	b	d
(2)	a	b	c	d
(3)	a	d	b	c
(4)	d	c	b	a

**Ans.** (1)

**Sol.** (i)  $Si(CH_3)_4$  Silane  
 (ii)  $Si(CH_3)_2(OH)_2$  Chain silicone  
 (iii)  $Si(CH_3)(OH)_3$  2D silicone  
 (iv)  $Si(CH_3)_3(OH)$  Dimeric silicone

10. 2 mole of  $NOCl$  taken in 1 L of closed container it dissociate into  $NO$  &  $Cl_2$  gas



At equilibrium if 0.4 mole of  $NO$  are obtained then value of  $K_c$  is  $\_\_\_\_ \times 10^{-4}$

**Ans.** (125)

**Sol.**  $2NOCl \rightleftharpoons 2NO + Cl_2$

Initially 2 mole 0 0  
 At equilibrium (2 - 0.4) 0.4 0.2  
 = 1.6

$$K_c = \frac{(NO)^2(Cl_2)}{(NOCl)^2} = \frac{(0.4)^2(0.2)}{(1.6)^2}$$

$$= 0.0125$$

$$= 125 \times 10^{-4}$$

**Ans.** = 125

11. The limiting molar conductivities of NaI, NaNO<sub>3</sub>, and AgNO<sub>3</sub> are 12.7 S m<sup>2</sup> mole<sup>-1</sup> and 13.3 S m<sup>2</sup> mole<sup>-1</sup> respectively (all at 25°C). The limiting molar conductivity of AgI at this temperature is :

Ans 14

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- Sol. Given  $\lambda_m^\infty(\text{NaI}) = 12.7 \text{ S m}^2 \text{ mole}^{-1}$   
 $\lambda_m^\infty(\text{AgNO}_3) = 13.3 \text{ S m}^2 \text{ mole}^{-1}$   
 $\lambda_m^\infty(\text{NaNO}_3) = 12 \text{ S m}^2 \text{ mole}^{-1}$   
 $\lambda_m^\infty(\text{AgI}) = \lambda_m^\infty(\text{AgNO}_3) + \lambda_m^\infty(\text{NaI}) - \lambda_m^\infty(\text{NaNO}_3)$   
 $= 13.3 + 12.7 - 12$   
 $= 14 \text{ S m}^2 \text{ mole}^{-1}$

12. Find % of Fe<sup>2+</sup> in Fe<sub>0.93</sub>O. [Report your answer to nearest integer]

Ans. (85)

- Sol.  $\text{Fe}_{0.93}\text{O}$
- $\begin{array}{cc} +2 & +3 \\ \swarrow & \searrow \\ x & (0.93-x) \end{array}$
- $2x + 3(0.93 - x) = 2$   
 $2x + 0.93 \times 3 - 3x = 2$   
 $2.79 - 2 = x$   
 $x = 0.79$   
 $\% \text{ of Fe}^{2+} = \frac{0.79}{0.93} \times 100 = 84.94\%$

Ans = 85

13. Identify correct match using Column I & Column II

Column I

Column II

(i) Spontaneous process

(a) Isothermal and isobaric process

(ii)  $\Delta H < 0$

(b)  $\Delta H < 0$

(iii)  $\Delta T = 0, \Delta P = 0$

(c)  $\Delta G < 0$

(iv) Exothermic process

(d) (Bond energy of reactant) – (Bond energy of product)

	I	II	III	IV	I	II	III	IV
(1)	c	d	a	b	(2)	b	a	c
(3)	d	b	c	d	(4)	a	d	b

Ans. (1)

- Sol. (i) For spontaneous process  $\Rightarrow \Delta G < 0$   
 (ii) For exothermic process  $\Rightarrow \Delta H < 0$   
 (iii) For isothermal process  $\Rightarrow \Delta T < 0$   
 For isobaric process  $\Rightarrow \Delta P < 0$

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14. Column I Column II

- |                                      |                     |
|--------------------------------------|---------------------|
| (i) Gold (Au)                        | (a) SO <sub>2</sub> |
| (ii) Blister Copper                  | (b) NaOH            |
| (iii) Al <sub>2</sub> O <sub>3</sub> | (c) NaCN            |
| (iv) Froth stabilizer                | (d) Aniline         |

Correct matching is

- |      |   |    |     |    |     |   |    |     |    |
|------|---|----|-----|----|-----|---|----|-----|----|
|      | I | II | III | IV |     | I | II | III | IV |
| (1*) | c | a  | b   | d  | (2) | a | b  | c   | d  |
| (3)  | b | c  | d   | a  | (4) | c | b  | a   | d  |

Sol. Theory based

15. Identify the correct products in following reaction



- (1) Be, LiAlCl<sub>4</sub>NCI (2) BeH<sub>2</sub>, LiCl, AlCl<sub>3</sub> (3) AlH<sub>3</sub>, BeH<sub>2</sub> HCl (4) Be, AlCl<sub>3</sub>, LiCl

Ans (2)

Sol.  $2\text{BeCl}_2 + \text{LiAlCl}_4 \rightarrow 2\text{BeH}_2 + \text{LiCl} + \text{AlCl}_3$

16. Column I Column I

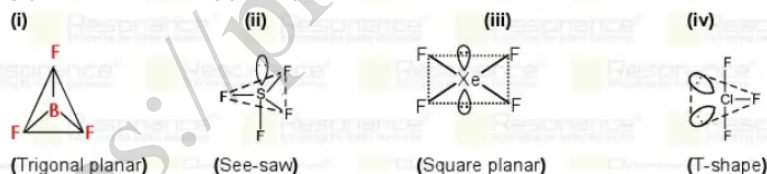
- |                        |                     |
|------------------------|---------------------|
| Compound               | Shape               |
| (i) BF <sub>3</sub>    | (a) T-Shape         |
| (ii) SF <sub>4</sub>   | (b) Square planar   |
| (iii) XeF <sub>4</sub> | (c) See - saw       |
| (iv) ClF <sub>3</sub>  | (d) Trigonal planar |

Correct matching is

- |     |   |    |     |    |     |   |    |     |    |
|-----|---|----|-----|----|-----|---|----|-----|----|
|     | I | II | III | IV |     | I | II | III | IV |
| (1) | d | c  | b   | a  | (2) | a | b  | c   | d  |
| (3) | d | a  | c   | d  | (4) | b | c  | d   | a  |

Ans (1)

Sol. (i) BF<sub>3</sub> (a) Trigonal planar  
 (ii) SF<sub>4</sub> (b) See - saw  
 (iii) XeF<sub>4</sub> (c) Square planar  
 (iv) ClF<sub>3</sub> (d) T-shape







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17. 0.76 gram mixture of H<sub>2</sub> and O<sub>2</sub> occupy 2000 cm<sup>3</sup> of volume at 300K and 10<sup>5</sup> Pascal pressure, then molar ratio of Hydrogen and oxygen is

Sol. For mixture.

$$PV = nRT$$

$$1 \times 2 = n_{\text{Total}} \times \frac{1}{12} \times 300$$

$$n_{\text{Total}} = \frac{24}{300} = 0.08 \text{ Mole}$$



$$n_{H_2} + n_{O_2} = 0.08$$

$$(i) W_{H_2} + W_{O_2} = 0.76$$

$$(ii) \frac{W_{H_2}}{2} + \frac{W_{O_2}}{32} = 0.08$$

$$16W_{H_2} + W_{O_2} = 0.08 \times 32 ;$$

$$W_{H_2} = \left( \frac{0.08 \times 32 - 0.76}{15} \right) = 0.12 ; \quad W_{O_2} = 0.64$$

$$\text{So } n_{H_2} = \frac{0.12}{2} = 0.06 ; \quad n_{O_2} = \frac{0.64}{32} = 0.02$$

Ans = 3

18. Uncertainty in position of a moving particle is  $10^{-7}$  m and uncertainty velocity is  $2.4 \times 10^{-24}$  m/sec, then mass of particle is  $[X] \times 10^{-6}$  Kg value of X is \_\_\_\_ . [Report your answer to nearest integer]

Sol. According to Heisenberg uncertainty Principle

$$\Delta x \times \Delta P \geq \frac{h}{4\pi}$$

$$\Rightarrow 10^{-7} \times m \Delta V = \frac{6.62 \times 10^{-34}}{4 \times 3.14}$$

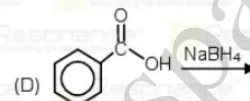
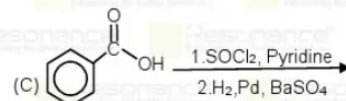
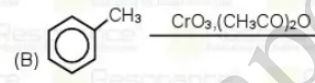
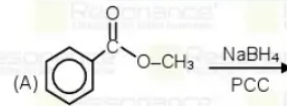
$$\Rightarrow 10^{-7} \times m \times 2.4 \times 10^{-24} = \frac{6.62 \times 10^{-34}}{4 \times 3.14}$$

$$M = 0.2196 \times 10^{-9} \text{Kg}$$

$$= 21.96 \times 10^{-6} \text{kg}$$

Ans = 22

19. Which of the following combination gives Benzaldehyde.



(1) A and B

(B) B and C

(C) A and D

(D) C and D





Ans. (2)

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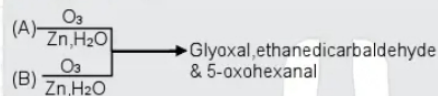
20. **Statement-I** In Hofmann rearrangement reaction alkyl group shift from carbonyl carbon atom to nitrogen atom.

**Statement-II** In alkyl shift, the group get shifted to the electron deficient carbon atom.

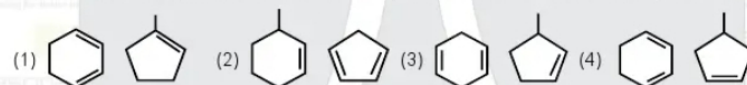
- (1) Both Statement-I & Statement-II are True. (2) Both Statement-I & Statement-II are False.  
(3) Statement-I is True while Statement-II is False. (4) Statement-I is False while Statement-II is True.

Ans. (1)

21. A mixture of organic compound A and B on reductive ozonolysis



Gives the product glyoxal, ethanedicarbaldehyde and 5-oxohexanal. The compound A and B are :



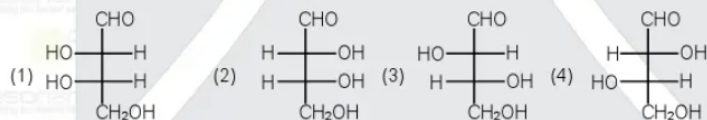
Ans. (1)

22. L-Isomer of an organic compound with molecular formula  $C_4H_8O_4$ , gives +ve Tollen's reagent test.

On treatment with bromine water gives optically active isomer but on treatment with concentrated  $H_2SO_4$  gives

optically inactive isomer. The correct figure of compound is :

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Ans. (1)

23. The correct match of the polymers given in Column-I with their uses in Column-II is

Column - I	Column - II
(A) Bakelite	(i) Water pipe
(B) Polyvinyl Chloride	(ii) Electric Switches
(C) Polystyrene	(iii) TV Cabinet
(D) Glyptal	(iv) Paint and Lacquers
(1) (A)-(ii), (B)-(i), (C)-(iv), (D)-(iii)	(2*) (A)-(ii), (B)-(i), (C)-(iii), (D)-(iv)
(3) (A)-(i), (B)-(ii), (C)-(iii), (D)-(iv)	(4) (A)-(ii), (B)-(iii), (C)-(i), (D)-(iv)





Ans. (2)

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24. **Statement-I:** Classical fog is reducing smog and found in cold and humid region.

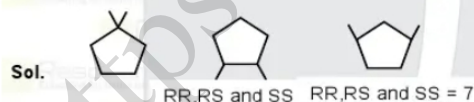
**Statement-II:** Photo chemical smog is oxidising and consist of oxidising agent as oxides of nitrogen nitrogen oxides, Volatile Organic Compounds (VOCs), tropospheric ozone, and PAN (peroxyacetyl nitrate)

- (1) Both Statement-I & Statement-II are True. (2) Both Statement-I & Statement-II are False.  
(3) Statement-I is True while Statement-II is False. (4) Statement-I is False while Statement-II is True.

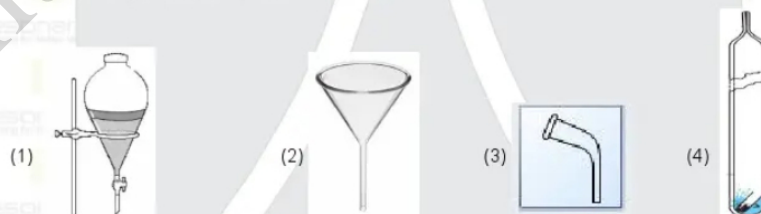
Ans. (1)

25. Total number of isomers of dimethylcyclopentane is (including stereoisomers)

Ans. (7)



26. Structure of separating funnel is

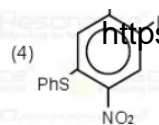
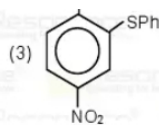
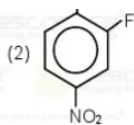
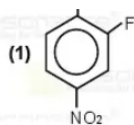


Ans. (1)

27. Major product of the given reaction is



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Ans. (2)

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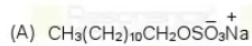
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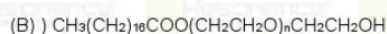
28. The correct match of the compounds given in Column – I with their uses in Column – II

Column – I

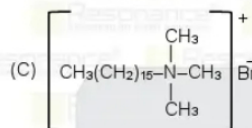
Column – II



(i) Toothpaste



(ii) Dish Washer



(iii) Hair conditioner



(iv) Soap

(1\*) (A)-(i), (B)-(ii), (C)-(iii), (D)-(iv)

(2) (A)-(ii), (B)-(iii), (C)-(iv), (D)-(i)

(3) (A)-(i), (B)-(iv), (C)-(iii), (D)-(ii)

(4) (A)-(iv), (B)-(iii), (C)-(i), (D)-(ii)

Ans. (1)

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